sults among one another will show whether any chlorine escapes or whether any other variable sources of error lie in the chemical reaction itself. The results of four such determinations follow. The weight of dish, cover and rod was less than 70 g.

Exp. A. Dish + $AgNO_3$ minus weight after = 0.5311 g. $w_1 - w = 0.0001$. Wt. sol. 50.0445 g.

Exp. B. Dish + $AgNO_3$ minus weight after = 0.5316 g. $w_1 - w = 0.0004$. Wt. sol. 50.0548 g.

Exp. C. Dish + $AgNO_3$ minus weight after = 0.53105 g. $w_1 - w = 0.0001$. Wt. sol. 50.0391 g.

Exp. D. Dish + $AgNO_3$ minus weight after = 0.5310 g. $w_1 - w = 0.00005$. Wt. sol. 50.0396 g.

Hence, the loss (corrected by companion dish) is respectively, per 50 g. of solution taken: in A 0.53063, in B 0.53062, in C 0.53095, in D 0.53054. Since the balance was sensitive to only 0.05 mg. it is evident that errors of weighing alone account fully for the surprisingly small variations observed. It is, further, clear that these errors may be diminished if there were any object in so doing, by operating with normal instead of 0.2 N solutions. The disadvantage that the loss of weight is a smaller quantity than the absolute weight of the HCl determined, so that all errors are multiplied by the factor 1.37, is made unimportant by the great accuracy of the process in itself. In spite of its precision, the method is not time-consuming, as the evaporations require no oversight. The time spent is, practically, only that required by the weighings. Most of this work was done in the laboratory of the Andrews Chemical Works, Davenport, Iowa.

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CORRECTION.

Correction for article on "The Potential of Silver in Nonaqueous Solutions of Silver Nitrate," by Vernette L. Gibbons and F. H. Getman, which appeared in the August number of THIS JOURNAL, page 1645: The last five values in Table X should have been headed Table XI.—Solutions in Pyridine.

[CONTRIBUTION FROM THE CHEMICAL LABORATORY OF THE UNIVERSITY OF WASHINGTON.]

BENZOYLATIONS IN ETHER SOLUTION.

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In a former paper[1] it was shown that acetyl chloride reacts in anhydrous ether solutions with primary, secondary and tertiary bases. The initial products were invariably additive products and these, by simple splitting or by hydrolysis, yielded the ordinary acetylated products.

[1] THIS JOURNAL, 34, 1399 (1912).